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In an alternate form, the Redlich-Kwong equation is written as  $Z = 1 / (1 - h) - (A/B) [h / (1 + h)]$  where  $h = b/V = B P / Z$ ,  $B =$

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$b/RT$ ,  $A/B = a/b RT$  1.5 , and  $Z$  ,  
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and  $Z_c$  are critical  
temperature, critical pressure,  
critical volume, and critical  
compressibility factor,  
respectively. Thus, for p-xylene,  $T_c =$   
 $412.3[0.567+0.106-(0.106)^2]-1$   
 $= 623.0\text{K}(661.8\text{ F})$ (literature  
value is 616.2 K)  $P_c =$   
 $106.16(0.34+1.378)^{-2} =$   
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